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The calculated answer (0.5639 mol) is indeed less than 1 mol, so we have probably not made a major error in the calculations. C To calculate the number of molecules in the sample, we multiply the number of moles by Avogadro ' s number: moleculesof ethylene glycol = 0.5639 mol(6.022 × 10²³ 1 mol) = 3.396 × 10²³ molecules.

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312 Chapter 11 The Mole Converting Number of Representative Particles to Moles
Zinc is used as a corrosion-resistant coating on iron and steel. It is also an essential trace element in your diet. Calculate the number of moles that contain 4.50×10^{24} atoms of zinc (Zn). 1. You are given the number of atoms of zinc and must find the equivalent number of moles.

Chapter 11: The Mole

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dividing the mass of the element in one mole of the compound by the mass of one mole of the compound and multiplying by 100. if there is more than one of the

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element in in the compound like CH_4 has 4 hydrogen, then you have to multiply the mass of hydrogen by 4 and then divide it by the mass of the CH_4 and then multiply by 100.

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